Gas review key from Chem. 120

1. A 6.75 L flask contains a fixed amount of gas at 31oC and a constant pressure. If the temperature is increased to 125oC, what will the volume of the gas be?

1. A vacuum pump exhausts a heavy-walled 1.50-L round-bottomed flask to a pressure of 3.55 x 10-6torr. How many particles are present if the temperature is 273 K?

How many molecules of xenon hexafluoride are required to react with 0.579 L of hydrogen gas at 2.46 atm and 45 °C in the unbalanced reaction show below (10 points)?

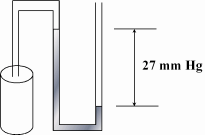
XeF6 (s) + 3 H2 (g) 🡪 Xe (s) + 6 HF (g)

1. It is found that 250. mL of an unknown ideal gas at STP has a mass of 2.50 g. What is the molar mass of the unknown gas?

1. A mixture of hydrogen, H2, nitrogen, N2, and Argon, Ar, gases is present in a steel cylinder. The total pressure within the cylinder is 675 mm Hg and partial pressures of nitrogen and argon gases are, respectively, 354 mm Hg and 235 mmHg. If carbon dioxide, CO2, gas is added to the mixture at constant temperature until the total pressure reaches 842 mm Hg, what is the partial pressure, in millimeters of Hg, of the following (8 points)?
   1. CO­2
   2. N2
   3. Ar
   4. H2
2. Nitrogen dioxide gas reacts with water vapor to produce oxygen and ammonia gases. Suppose that 12.8 g of nitrogen dioxide reacts with 5.00 L of water vapor at 375 °C and 725 torr (24 points).
3. Write the balanced chemical equation
4. Use an ICE table to find the limiting reagent and complete the table
5. How many grams of ammonia are produced?
6. What is the percent yield if 0.984 g NH3 are collected?
7. How many molecules of the excess reagent remain?

Newer material

1. What is the pressure, in atmospheres, of the gas in the following open-tube manometer? Assume that atmospheric pressure is 102 kPa. (5 points)



1. An artificial atmosphere for the planet Jupiter consists of 92% hydrogen and 8% helium. 15g of this gas mixture are placed in a 50L container at –25°C. What is the partial pressure of each gas? (5 points) What is the mole fraction of hydrogen and helium in the artificial atmosphere?
2. **SHOW ALL WORK.** You are asked to select a high precision valve that will be used to smoothly deliver uranium hexafluoride (UF6) at a rate of 0.025 L/min. For safety and economic reasons, you decide to use helium (He) to test the flow rate of the new valve. Determine the correct He flow rate of this valve in L/min.
3. For a spacecraft or a molecule to leave the moon, it must reach the escape velocity (speed) of the moon, which is 2.37 km/s. The average daytime temperature of the moons surface is 365 K. What is the rms speed (in m/s) of a hydrogen molecule at this temperature? How does this compare with the escape velocity?
4. A given volume of nitrogen, N2, required 68.3 s to effuse from a hole in a chamber. Under the same conditions, another gas required 85.6 s for the same volume to effuse. What is the molecular weight of this gas?
5. Calculate the pressure of water vapor at 120.0° C if 1.000 mol of water vapor occupies 32.50 L. Use the van der Waals equation (see Table 5.7 for data). Compare the result from the ideal gas law.  (a= 5.537 L2-atm/mol2 and b = 0.03049 L/mol)